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Molecular Geometry 3 9. Explain the difference between a bonding electron domain and a nonbonding electron domain using the examples in Model 1. 10. Circle the correct word or phrase to complete the sentences: Pairs of electrons will (attract/repel) each other.

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Molecular Geometry 3 9. Explain the difference between a bonding electron domain and a nonbonding electron domain using the examples in Model 1. 10. Circle the correct word or phrase to complete the sentences: Pairs of electrons will (attract/repel) each other.

### **20 Molecular Geometry-S**

following molecular shapes: Trigonal planar AX E Bent  $< 109.50$   
Notation Geometry Electron group arrangement Molecular Shape

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Idealized Bond Angles Linear 180° Bent (V-shaped) < 120°  
Trigonal pyramidal < 109.5° Tetrahedral 109.5°  
17. Draw the Lewis Dot structure of 1-12 IL) 18.

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Molecular Geometry How can molecular shapes be predicted using the VSEPR theory? Why? When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The true shape of a molecule is important because it determines many physical and chemical properties for the substance.

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13, Based on the information in the Read This! section, sketch one of the molecular shapes shown below in each of the boxes provided in Model I. Trigonal planar 120° Bent 104.5° Three-Dimensional Molecular Shapes Tetrahedral 109.5° STOp Molecular Geometry Linear Pyramidal 107°

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consider the external and internal pressure data given in Model  
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